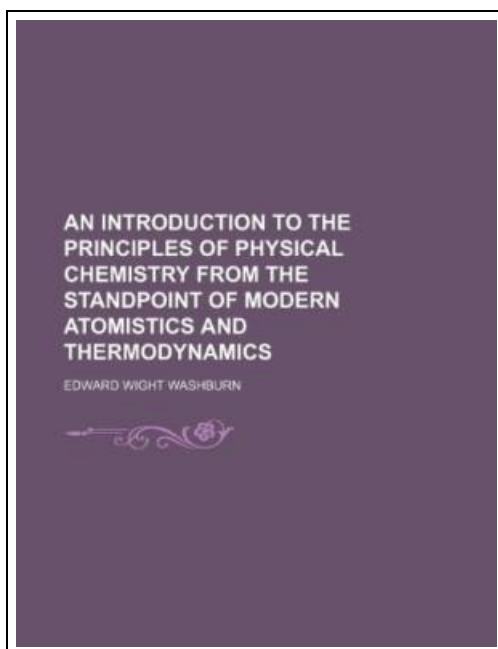


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AN INTRODUCTION TO THE PRINCIPLES OF PHYSICAL CHEMISTRY FROM THE STANDPOINT OF MODERN ATOMISTICS AND THERMODYNAMICS



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1915 Excerpt: .is known as Faraday s Law of Electrolysis and was discovered by Michael Faraday in 1833. The proportionality constant, F , in equation (23) is known as one faraday or one equivalent of electricity and is evidently the amount of electricity required to produce one equivalent of chemical change. The quantities, e and F (equation 22), evidently bear the same relation to each other that the weight of an atom does to the atomic weight of the element (l , 7). Michael Faraday (1791-1867). The son of a blacksmith in Surrey, England. Served as laboratory assistant to Sir Humphrey Davy and later became Professor of Chemistry at the Royal Institution, London. He discovered benzene in 1825 and made a large number of important discoveries in electricity and magnetism. The experiments upon which he based his law of electrolysis are an interesting illustration of the keen insight which led Faraday to enunciate a general law upon what seems today to be very meagre and inaccurate data. (Faraday s Experimental Researches in Electricity. Everyman s Library Series, pp. 48 to 84.) The value of the f araday has been very accurately determined⁶ by measuring (with a silver coulometer) the quantity of electricity required to deposit one equivalent weight of metallic silver from a silver nitrate solution and also by measuring (with an iodine coulometer) the amount required to liberate one equivalent weight of iodine from a solution of potassium iodide. The result obtained was $F = 96,500 (+ 0.01 \text{ per} \dots$

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