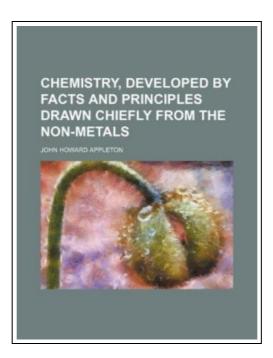
# Chemistry, Developed by Facts and Principles Drawn Chiefly from the Non-Metals



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\* This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1884 Excerpt: .of Carbon, Oxygen, Carbon dioxide, 12 32 44 Parts by weight. Parts by weight. Paris by weight. And likewise when iron is burned, there is formed an oxide Wiiosmj composition is expressed by the formula, Fe304; (to this substance the chemical name ferroso-ferric oxide is applied). So when sulphur is burned, sulphur dioxide is formed (SOs). 127 In this case the chemical change is represented by the following equation: Compound of Oxygen with Hydrogen. It haa already been shown that the hydrogen escaping from a suitable tube may be lighted in the air. If the burning jet is introduced into Fig. 36.--Hydrogen gas, generated by use of zinc and sulphuric acid, is then passed through a drying tube containing calcic chloride (CaCl ). By the act of combustion the union of the dried gas with oxygen of the air produces drops of water. oxygen gas the same combustion proceeds, only with greater energy. In either case there is produced a compound of hydrogen and oxygen. This compound is represented by the formula H2O, a formula epreeenting no other than the familiar substance water. At the moment of combustion of hydrogen very great heat is generated. In fact, a pound of hydrogen, upon burning in pure oxygen, yields about four times as much heat as a pound of pure carbon does in burning under the same favorable conditions. Indeed, the pound of hydrogen, when in combustion, yields more heat than a pound of any other substance known....

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